# Role of Oxidants in Microbial Pathophysiology

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### INTRODUCTION

Reactive oxygen species have been increasingly implicated as playing a central role in the pathophysiology of clinical infections. More specifically, superoxide, hydrogen peroxide, hydroxyl radical, hypohalous acid, and recently, nitric oxide are thought to contribute to these processes. These compounds exhibit a broad spectrum of biotoxicity and are crucial to host defense for the optimal microbicidal activity of neutrophils and other phagocytes (148, 209, 216, 313). In response, microorganisms have developed complex strategies not only to avoid contact with phagocyte-derived oxidants but also to defend themselves from injury once oxidants are encountered. Host cells have developed similar adaptations to protect themselves against a deleterious consequence of oxidant exposure, inflammatory tissue injury (209, 313). This review will discuss the formation of oxidants in vivo and their central role in the complex interplay between microbial invasion and host defense.

# GENERATION AND TOXICITY OF SELECTED OXIDANTS IN BIOLOGIC SYSTEMS

Many biochemical reactions vital to normal aerobic metabolism of human and microbial cells require the transfer of four electrons to molecular oxygen to form H<sub>2</sub>O. Under most circumstances, this transfer occurs simultaneously without the formation of other intermediates. However, molecular oxygen does have the capacity to undergo sequential univalent reduc-

tion to form other oxygen intermediates with different toxicities prior to the generation of  $H_2O$ .

The addition of one electron to  $O_2$  yields the superoxide radical  $(O_2^-)$ , which at physiologic pH rapidly reduces itself (dismutes,  $k \approx 2 \times 10^5 \ \mathrm{M}^{-1} \ \mathrm{s}^{-1}$ ) to form the divalent oxygen reduction product, hydrogen peroxide  $(H_2O_2)$ . Trivalent oxygen reduction in vitro occurs via the reaction of  $H_2O_2$  with  $O_2^-$  to produce the hydroxyl radical (OH). However, at physiologic pH, this reaction is of little biologic importance unless a transition metal catalyst (e.g.,  $\mathrm{Fe}^{3+}$ ) is present to enhance the reaction rate, yielding OH via the Haber-Weiss reaction (123) (Table 1). As discussed below, not all iron complexes can serve as a catalyst in this reaction (125). Besides OH formation, experimentally induced interactions between  $H_2O_2$  and iron chelates may also lead to the production of the reactive iron peroxocomplex and ferryl ion (268, 321). However, their role in human and microbial physiology is largely unknown.

Although most investigations have focused on 'OH formation via the Haber-Weiss mechanism, evidence also exists for the formation of 'OH from 'O<sub>2</sub>'-mediated reduction of hypochlorous acid (HOCl) (51, 189, 232, 250). A potent oxidant in itself, HOCl is generated by the interaction of H<sub>2</sub>O<sub>2</sub> with phagocyte-derived peroxidases (148).

Recently, intense investigation has been directed at another oxidant species, nitric oxide (NO'). NO is not a classic product of O<sub>2</sub> reduction; instead, its formation in mammalian cells is dependent on a group of enzymes termed nitric oxide synthases (NOS) (216, 224). These enzymes oxidize L-arginine to L-citrulline and NO. Although several related NOS isoforms have been isolated, they are divided into two categories, constitutive and inducible, based on differences in regulation and activities. Constitutive isoforms (cNOS) are found in neuronal

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TABLE 1. Chemical reactions involving reactive oxygen species

Reaction	Formula <sup>a</sup>
Haber-Weiss reaction	$O_2^- + Fe^{3+} \rightarrow O_2 + Fe^{2+}$ $H_2O_2 + Fe^{2+} \rightarrow OH + OH^- + Fe^{3+}$ $O_2^- + H_2O_2 \rightarrow OH + OH^- + O_2$
Myelo (eosinophil) peroxidase	$H_2O_2 + HX \rightarrow HOX + H_2O$
Nitric oxide synthase	L-Arginine → L-citrulline + NO
Peroxynitrite formation/ decomposition	NO' $+$ 'O <sub>2</sub> <sup>-</sup> $\rightarrow$ ONOO <sup>-</sup> ONOO <sup>-</sup> $+$ H <sup>+</sup> $\rightarrow$ ONOOH ONOOH $\rightarrow$ 'OH $+$ NO'
	2 GSH + $H_2O_2 \rightarrow$ G-S-S-G + $2H_2O$ G-S-S-G + $2NADPH \rightarrow 2GSH + 2NADP$
Catalase	$2H_2O_2 \rightarrow 2H_2O + O_2$
SOD	$2^{\circ}O_2^{-} + 2H^+ \rightarrow H_2O_2 + O_2$

a X. halide.

and endothelial cells. cNOS activity responds to changes in intracellular calcium concentration via calcium-calmodulin binding. This results in the intermittent production of small amounts of NO necessary for physiologic processes such as neurotransmission and blood pressure regulation. By using spin-trapping techniques, brain NOS has also been demonstrated to generate  $O_2^-$  in a calcium-calmodulin-dependent manner (243). The inducible NOS isoform (iNOS) is expressed in many cell types, including hepatocytes, respiratory epithelium, and macrophages. Its activity is independent of fluctuations in the intracellular calcium concentration. Factors known to modulate iNOS levels include a number of cytokines, microorganisms, and microbial products, consistent with the importance of iNOS activity in host defense and inflammation. Many bacterial species are also capable of generating NO under conditions of low oxygen tension via nitrite reductases (334). Once formed, NO has the ability to act as an oxidizing agent alone or interact with  $O_2^-$  to generate peroxynitrite (ONOO<sup>-</sup>) ( $k \approx 6.7 \times 10^9 \text{ M}^{-1} \text{ s}^{-1}$ ) (245, 247, 284) and ultimately OH via peroxynitrate formation and decomposition (Table 1) (19). Although a transition metal catalyst is not required in this system, thermodynamic and kinetic considerations may not favor the formation of 'OH via this reaction in vivo (170).

### Superoxide and Hydrogen Peroxide

Superoxide is a moderately reactive compound capable of acting as an oxidant or reductant in biologic systems. This relative inactivity allows  $O_2^-$  to diffuse for considerable distances before it exerts its toxic effects. Extracellularly generated  $O_2^-$  can gain access to intracellular targets via cellular anion channels (264). These targets include bacterial enzymes, particularly those involved in biosynthesis of branched amino acids (e.g.,  $\alpha,\beta$ -dihydroxyisovalerate dehydratase and NADH-bound lactic dehydrogenase) (122, 174). Several *Escherichia coli* (and mammalian) dehydratases containing [4Fe-4S] clusters are particularly susceptible to inactivation by  $O_2^-$ , including aconitase, 6-phosphogluconate dehydratase,  $\alpha,\beta$ -dihydroxyacid dehydratase, and fumarases A and B (101, 108, 109, 111, 183). Aconitase has also been shown to be inactivated by

ONOO<sup>-</sup> but not NO (57, 139). These enzymes are unique in that they can subsequently undergo reactivation by an iron-dependent mechanism (110). It is postulated that the inactivation occurs at an early stage of oxidative stress, such that aconitases function as "circuit breakers," halting the production of toxic 'O<sub>2</sub><sup>-</sup> by temporarily shutting down cellular oxidative metabolism (109). Once the stress has passed, the dehydratases can be reactivated by intracellular iron and thiols rather than having to be synthesized de novo (109).

In environments of low pH, such as at sites of inflammation or inside the phagosome,  $O_2^-$  becomes protonated to form  $HO_2$ . Because of its neutral charge,  $HO_2$  is more membrane permeable and more likely to react with itself to form  $H_2O_2$ . Additional toxicity of  $O_2^-$  in biologic systems is likely to occur via its participation in the Haber-Weiss reaction in the presence of catalytically active iron (123).

Hydrogen peroxide is a more reactive oxidant than  $O_2^-$ , and readily diffuses across cell membranes. Potential sources of H<sub>2</sub>O<sub>2</sub>-mediated damage of cellular constituents include the oxidation of cellular membranes and enzymes, DNA damage and mutagenesis, and the inhibition of membrane transport processes (313). Imlay and Linn have described in greater detail the mechanisms of H<sub>2</sub>O<sub>2</sub>-mediated damage. They demonstrated that killing of E. coli by H<sub>2</sub>O<sub>2</sub> is bimodal in that low (1 to 3 mM) and high (>20 mM) concentrations of  $H_2O_2$  are more lethal than intermediate concentrations (153). Mode 1 (low H<sub>2</sub>O<sub>2</sub> concentration) killing has been attributed to DNA damage mediated by the interaction of H<sub>2</sub>O<sub>2</sub> with Fe<sup>2+</sup> to form the toxic ferryl radical (149), an intermediate product in the formation of OH. Exposure of E. coli to these low concentrations of H<sub>2</sub>O<sub>2</sub> induces a protective response which confers increased resistance to subsequent H<sub>2</sub>O<sub>2</sub> exposures by an enhanced ability to carry out recombinational DNA repair (154). Mode 2 killing, which does not require iron or an electron source to occur, is not due to DNA damage but may involve the oxidation of a separate cellular target (137). Recent data by Pacelli et al. demonstrate that NO potentiates H2O2induced killing of E. coli (235). This suggests that macrophagederived NO, in addition to its own cytotoxic effects, may interact with H<sub>2</sub>O<sub>2</sub> to enhance microbicidal activity at sites of infection (235).

## **Hydroxyl Radical**

In many cases where  ${\rm O_2}^-$  and/or  ${\rm H_2O_2}$  is implicated in cell injury, it is unclear whether the process is mediated by these compounds or whether they simply serve as precursors for another, more potent oxidant species (e.g., OH), which is truly mediating the injury. Studies using more sensitive free radical detection systems implicate OH in the oxidation of a large number of biomolecules including proteins, DNA, and lipids, as a result of their initial exposure to  ${\rm O_2}^-$  and/or  ${\rm H_2O_2}$ . Owing to its high reactivity, OH is diffusion limited such that once formed in a biologic system, it is likely to travel only very short distances before it encounters an oxidizable substrate. This property dictates that OH must be generated in close proximity to a critical cellular target molecule in order for it to mediate injury directly (77).

A mechanism by which OH and other oxidants may cause cell injury at sites distant from their formation is via the initiation of a free radical cascade (43, 313). Oxidation of unsaturated fatty acids within a lipid membrane can produce peroxyl radical, which in turn can react with other nearby lipid molecules to generate additional lipid radicals. These new lipid radicals can then react with other unsaturated lipids, thereby setting up a free radical chain reaction (43). This reaction

eventually results in the oxidation of biomolecules at sites considerably distant from where the initial free radical reaction occurred (43).

### Sources of Iron Available for Hydroxyl Radical Generation In Vivo

Since OH formation from  $O_2^-$  and  $H_2O_2$  under physiologic conditions requires the presence of a transition metal catalyst, there has been considerable interest in determining which iron chelates potentially present in vivo could serve as 'OH catalysts. In humans, intracellular iron is predominantly complexed to ferritin in a relatively noncatalytic form (301). Likewise, almost all extracellular iron is tightly bound to host binding proteins (transferrin and lactoferrin) in forms unable to catalyze 'OH formation (7, 9, 42, 45, 320). In fact, there are strong data suggesting that lactoferrin serves as an antioxidant (34, 39, 40, 71, 215). Neutrophil lactoferrin may function to trap iron from ingested microorganisms (215). In phagocytes that do not contain lactoferrin (i.e., monocytes and macrophages), a specific surface receptor binds exogenous lactoferrin (20, 24, 50, 213, 326). Monocytes/macrophages previously incubated with lactoferrin are less susceptible to iron-dependent peroxidation of their membranes (40). Thus, via its interaction with phagocytes, lactoferrin may prevent iron-catalyzed oxidant formation, thereby limiting inflammatory tissue injury. This would complement the ability of the protein to limit the availability of iron for microbial growth (88, 99). Lactoferrin also binds lipopolysaccharide, an important compound mediating toxicity in sepsis. Although this interaction has no effect on the ability of lactoferrin to inhibit the Haber-Weiss reaction, it does disrupt lipopolysaccharide priming of phagocytes for 'O<sub>2</sub><sup>-</sup> production

In contrast, in vitro data suggest that modification of some host iron-chelating proteins by proteases or  $O_2^-$  can result in the generation of products capable of catalyzing 'OH formation (23, 33). The Pseudomonas aeruginosa secretory product Pseudomonas elastase and other host-derived proteases present at sites of inflammation are known to cleave transferrin and lactoferrin into lower-molecular-weight iron chelates (27, 32, 84, 90, 91, 254). Pseudomonas elastase-cleaved transferrin and, to a lesser extent, lactoferrin are capable of catalyzing 'OH formation when a source of 'O<sub>2</sub> and H<sub>2</sub>O<sub>2</sub> is concurrently present (33, 208, 211). Additional studies have demonstrated the ability of Pseudomonas elastase and other protease-cleaved transferrin to enhance oxidant-mediated porcine pulmonary artery endothelial cell injury via 'OH generation in an in vitro model (208). Evidence supporting the potential clinical relevance of these findings has been obtained by the detection of transferrin cleavage products in bronchoalveolar lavage specimens from P. aeruginosa-infected cystic fibrosis patients but not in those from normal individuals (35).

As microorganisms require iron for growth and replication, their mechanisms of iron acquisition and storage have evolved to fulfill these needs. Intracellular bacterial iron is primarily complexed to ferritin-like iron storage proteins (229, 290). To acquire iron from the extracellular environment, aerobic and facultative anaerobic bacteria, as well as fungi, synthesize diverse low-molecular-weight Fe(III)-scavenging ligands collectively termed siderophores (229). These compounds possess a high affinity for iron, which is probably important at sites of infection (e.g., the airway), where the availability of iron for bacteria is extremely limited due to competition from host iron-binding proteins. As an example, to be able to compete for iron effectively, *P. aeruginosa* synthesizes and secretes two types of siderophores: pyochelin and pyoverdin (74, 75). Stud-

ies from our laboratory suggest that pyochelin may play an important role not only in iron acquisition but also in P. aeruginosa-associated inflammatory tissue injury (38). The environment at sites of P. aeruginosa infections is replete with 'O<sub>2</sub>and H<sub>2</sub>O<sub>2</sub> generated by local phagocytes and via the redox action of pyocyanin on target cells (discussed below). Ferripyochelin can act as a catalytically active iron chelate in the formation of 'OH (37) and can enhance oxidant-mediated in vitro porcine pulmonary artery endothelial (38) and epithelial (36) cell injury. Therefore, although the production of siderophores by P. aeruginosa is an adaptive mechanism for obtaining necessary iron under stressful conditions, the same compounds may also potentiate oxidant-mediated tissue injury via the catalysis of OH. A similar role for pyoverdin has not been found (69). Nevertheless, it is possible that siderophores produced by other organisms play a similar role, but there are currently no available data to substantiate this.

Additional potential sources of catalytically active iron related to host-microbe interactions in vivo include iron released from hemoglobin through the action of the bacterial toxin hemolysin, host cell exposure to bacterially derived iron reduction compounds such as pyocyanin (discussed in a later section), and/or the release of intracellular iron from damaged mammalian or bacterial cells into the microenvironment. Regardless of their source, it is necessary that extracellularly generated iron catalysts remain in close proximity to the cell in order to facilitate OH-mediated injury, given the limited diffusibility of OH (116).

#### **Myeloperoxidase-Derived Oxidants**

Coincident with their production of  $O_2^-$  and  $H_2O_2$ , stimulated human phagocytes release one of two distinct peroxidases from their cytoplasmic granules. In the case of neutrophils and monocytes, this enzyme is myeloperoxidase (MPO), whereas for eosinophils it is eosinophil peroxidase (EPO) (148). The interaction of MPO and EPO with H<sub>2</sub>O<sub>2</sub> forms hypohalous acids (HOX, where X = halide). It is generally thought that macrophages lack either enzyme (169); however, recent data suggest that this may not be universally true (79). Myeloperoxidase is a glycoprotein (molecular weight, 150) consisting of a pair of glycosylated heavy ( $\alpha$ )-light ( $\beta$ ) protomers, each of which contains an iron atom (225). EPO is an  $\alpha\beta$  glycoprotein, similar in structure to hemi-MPO (148). These enzymes are cationic, thus allowing them to stick to cell surfaces and perhaps enhancing their potential for cell injury by increasing the local concentration of hypohalous acid at the target cell membrane (184, 212, 251, 277, 325).

Hypohalous acids are potent oxidants known to have several cytotoxic effects on mammalian and bacterial cells. Cell membrane integrity may be violated by membrane peroxidation and the oxidation and/or decarboxylation of membrane proteins (2, 322). Likewise, oxidation of components of the bacterial respiratory chain and interference with bacterial DNA-membrane interaction required for bacterial division can disrupt normal cellular metabolism and replication (249, 265). Activated neutrophils and monocytes can also generate cytotoxic chloramines, tyrosyl radical, and OH via an MPO-dependent pathway (140, 148, 314).

# Nitric Oxide

Nitric oxide is cytostatic or cytotoxic for both prokaryotic and eukaryotic cells (105, 216). The primary mechanism of injury involves the interaction of NO with iron-containing moieties in key enzymes of the respiratory cycle (e.g., glyceraldehyde-3-phosphate dehydrogenase) and with DNA synthesis

leading to mutagenesis in target cells (216). Nitric oxide also can react with other biomolecules to form new compounds that are also capable of toxicity. For example, the formation of nitrosothiol groups on proteins can lead to the inactivation of enzymes or changes in protein function (216, 261). These groups can react further to cross-link sulfhydryl groups and thus initiate a chain reaction (261). In addition, NO and its derivatives can form toxic alkylating agents by reacting with secondary amines (156).

As shown in Table 1, ONOO<sup>-</sup> is generated by the reaction of O<sub>2</sub><sup>-</sup> with NO (19). Its ability to directly oxidize sulfhydryl groups and DNA bases, catalyze iron-independent membrane lipid peroxidation, and react with metals or metalloproteins (e.g., superoxide dismutase [SOD]) to form the toxic nitronium ion (NO<sup>2-</sup>) has led some investigators to suggest that ONOO<sup>-</sup> plays a more important role than its precursor NO in mediating cytotoxicity (19, 144, 155, 247, 248, 287). In addition, evidence suggests that upon protonation, ONOO<sup>-</sup> can undergo homolytic cleavage to form 'OH by an iron-independent mechanism (19); however, the biologic relevance of this reaction has not been definitively addressed.

# SOURCES OF OXIDANTS ENCOUNTERED BY MICROBES IN VIVO

### **Endogenous Sources**

Like eukaryotic cells, aerobic microorganisms are continually exposed to endogenous sources of toxic oxygen species as a consequence of aerobic metabolism (12). As discussed above, this occurs by sequential univalent electron reduction of molecular  $O_2$  to generate such species as  $O_2^-$ ,  $H_2O_2$ , and OH. Under certain conditions, homolytic cleavage of  $H_2O_2$  may also yield OH. These toxic oxygen species also can be generated as by-products of reactions involving glucose oxidase, xanthine oxidase, and thiol groups and flavins (73, 112, 252, 253). Furthermore, microbial exposure to UV or  $\gamma$  irradiation induces  $O_2^-$  production (200). Anaerobic organisms are particularly susceptible to oxidants derived via the above mechanisms, as they often lack the antioxidant defense mechanisms observed in aerobic organisms (discussed below).

A number of microorganisms, including Enterococcus faecalis (308), E. coli (150), Lactobacillus spp. (316), Streptococcus pneumoniae (316), and a number of Mycoplasma spp. (192, 204), also generate extracellular  $O_2^-$  and  $H_2O_2$ . Additional studies have shown that these oxidants can exert a number of beneficial and toxic effects on both the host and other microorganisms. For example, H<sub>2</sub>O<sub>2</sub>-producing Lactobacillus spp. inhibit Neisseria gonorrhoeae and human immunodeficiency virus (HIV) replication in vitro (167, 332), suggesting a nonspecific antimicrobial defense mechanism resulting from the presence of lactobacilli in the normal vaginal flora. Likewise, in women with bacterial vaginosis, H<sub>2</sub>O<sub>2</sub>-producing lactobacilli are notably absent from the vaginal flora (89). In contrast,  $O_2^$ made by Mycoplasma pneumoniae can inactivate host cell catalase, resulting in progressive oxidative damage to infected cells in vitro (3). S. pneumoniae-derived H<sub>2</sub>O<sub>2</sub> may play a role in host cellular injury in pneumococcal pneumonia, as it has been shown to be toxic to rat alveolar epithelial cells in an in vitro model (86). The formation of dental plaque and the subsequent development of gingivitis and periodontitis are also related to the balance of H<sub>2</sub>O<sub>2</sub>-producing and H<sub>2</sub>O<sub>2</sub>-degrading organisms in the oral microenvironment (269).

Microorganisms are also continually exposed to endogenously produced NO through denitrification (334). This process is a distinctive mode of respiration that is essential to

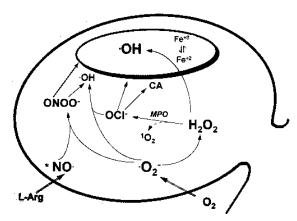


FIG. 1. Overview of intraphagosomal processes leading to oxidant-mediated microbial killings. CA, chloramines. Note that nitric oxide production occurs only in phagocytes with an inducible nitric oxide synthase. Reproduced from reference 209 with permission of the publisher.

many forms of bacteria and fungi; it involves the transformation of oxyanions of nitrogen to N<sub>2</sub>, mainly under conditions of reduced oxygen tension or strict anaerobiosis. Denitrification is controlled by a number of metalloenzymes, of which nitrite reductase has been identified as the enzyme responsible for the conversion of nitrite to NO (334). Two mutually exclusive nitrite reductases have been identified among denitrifying bacteria: a tetraheme cytochrome cd located in the periplasm of gram-negative organisms, and a Cu-containing protein bound to the cytoplasmic membrane of gram-positive organisms (334). The locations of these enzymes may limit the potential toxicity of the endogenously produced NO, as NO is subsequently rapidly reduced by a cytoplasmic membrane-associated NO reductase. Recent data also demonstrate the existence of a NOS system in Nocardia spp. (64), the first confirmation of such a system in microorganisms. The stoichiometry of products formed with respect to substrates used, cofactor requirements, and inhibition by  $N^{G}$ -nitro-L-arginine were found to be similar to those observed in mammalian NOS (64). There is evidence that erythrocytes infected with Plasmodium falciparum may also generate NO via NOS and produce a soluble factor that is able to evoke NO production in host tissues (113). Reports of other microbial NOS are likely to appear in the future.

## **Exogenous Sources**

#### Phagocyte-derived oxidants and their role in host defense.

The primary source of exogenous oxidative stress for pathogenic bacteria during the process of active infection is their attack by host phagocytic cells. Phagocytes utilize the cytotoxic effects of many of the oxidants outlined above as a component of their host defense mechanism (Fig. 1). When a phagocyte encounters a microorganism, the latter is surrounded by a portion of the phagocyte membrane, which then invaginates, forming a discrete phagosome (148). This process leads to increased phagocyte oxygen consumption and initiates a complex biochemical signaling system which activates a unique membrane-associated NADPH-dependent oxidase complex (67). This enzyme univalently reduces O<sub>2</sub> to O<sub>2</sub><sup>-</sup>, which is then secreted into the phagosome (67). There, O<sub>2</sub><sup>-</sup> dismutes to H<sub>2</sub>O<sub>2</sub>. These toxic compounds may also leak extracellularly as the phagosome is closing.

Following phagocytosis, microorganisms are subjected to

further insult as phagocyte primary (or azurophilic) cytoplasmic granules fuse with the phagosome. In addition to MPO, these granules contain mainly hydrolases (acid hydrolases, lysozyme, neutral proteases, deoxyribonucleases, etc.), which are probably responsible for the decomposition of killed organisms (31). Secondary (or specific) cytoplasmic granules fuse with the external plasma membrane before the primary granules do, thereby secreting their contents (lactoferrin, lysozyme, and vitamin  $B_{12}$ -binding protein) extracellularly (283). The membranes of these secondary granules also contain a number of functionally significant proteins, including CD11b/CD18, the formyl-methionyl-leucyl-phenylalanine receptor, and cytochrome  $b_{558}$  (48). The fusion of these granules with the plasma membrane serves to reinforce or sustain various cellular responses (29).

The importance of the NADPH-oxidase system for host microbicidal activity is exemplified in individuals with chronic granulomatous disease (CGD), a group of inherited disorders which are each characterized by defects in the NADPH-oxidase complex resulting in a lack of phagocyte 'O2 production (299). The NADPH-oxidase requires the assembly of its membrane and cytosolic components for generation of the respiratory burst. Likewise, the genetic defects in this enzyme observed among CGD patients are characterized by their localization to the membrane or cytosol. Approximately 60% of CGD patients have an X-linked defect in the membrane b cytochrome component as a result of mutations in the gp91phox (55%) or p22phox (5%) gene encoding the large and small subunits, respectively (236, 267, 300). Patients with autosomal defects most commonly lack the cytosolic component, p47*phox*, and account for approximately 35% of cases (55, 68). Less than 5% of CGD patients lack the p67phox cytosolic component (68). Regardless of the location of the defect, the clinical manifestations of the different genetic forms of CGD are quite similar. These persons suffer from recurrent pyogenic infections with organisms that are normally rapidly killed by oxidants: Staphylococcus aureus, enteric gram-negative rods, Aspergillus spp., and Candida spp. Infectious complications, which can involve virtually any organ system, typically begin in infancy and recur throughout childhood and adolescence.

Although associated with states of neutropenia, infections with other pathogenic organisms such as P. aeruginosa are infrequently encountered in CGD patients (299). In vitro data have demonstrated the ability of neutrophils to destroy P. aeruginosa (127, 145, 196, 221, 328). This process is markedly enhanced in the presence of serum which opsonizes the organism with complement and immunoglobulin to promote more efficient phagocytosis by the neutrophil (328). However, additional in vitro observations suggest that oxidants are not critical for neutrophil-mediated killing of *P. aeruginosa*. Neutrophils from CGD patients have the same capacity as normal neutrophils to kill P. aeruginosa (145). Without the presence of ambient  $O_2$ , neutrophils are unable to generate  $O_2$  and  $O_2$ . However, their abilities to kill P. aeruginosa under aerobic and anaerobic conditions appear to be similar (196). This is in contrast to findings with S. aureus, in which neutrophils are unable to kill the organism under anaerobic conditions in vitro (197). It may be that 'O<sub>2</sub>--independent mechanisms of neutrophil killing, such as those involving granule-derived proteins and proteases, are more important in *P. aeruginosa* elimination (311). However, these findings do not eliminate the possibility that neutrophil-derived oxidants increase the effectiveness of the O<sub>2</sub>-independent killing mechanisms.

Phagocyte-derived  $H_2O_2$  may also be converted intra- or extracellularly to HOCl and the longer-lived chloramines in the presence of chloride and myeloperoxidase. In addition,

MPO can catalyze the reaction of 'O<sub>2</sub><sup>-</sup> and HOCl to form 'OH (232). All of these compounds are known to have a number of cytotoxic effects in vitro (313). However, their overall significance in in vivo microbicidal activity is unclear, as patients with MPO deficiency demonstrate delayed killing of fungi and bacteria but are normally resistant to most infections (237). Of all patients recognized with this disorder (a prevalence of approximately 1 in 2,000 of the general population) (225), only a few have had serious infections (58, 166). The majority of these patients had visceral or disseminated candidiasis (58, 179). Three of these patients had concomitant diabetes mellitus (58, 179), perhaps indicating that the clinical morbidity associated with MPO deficiency requires an additional defect in host defense.

Phagocytes may also participate in mode 1 and 2 bacterial killing by generating  $H_2O_2$  as described by Imlay et al. (149, 153). As discussed above, the interaction of exogenous  $H_2O_2$  at low concentrations with intracellular Fe<sup>2+</sup> in *E. coli* results in DNA damage mediated by the ferryl radical. Bacterial exposure to higher  $H_2O_2$  concentrations resulted in killing by a separate oxidative mechanism.

Perhaps a more physiologically significant mechanism involved in phagocyte-mediated oxidant generation and microbial toxicity involves the iron-catalyzed intra- or extracellular reaction of 'O<sub>2</sub> and H<sub>2</sub>O<sub>2</sub> to form 'OH. Although there is a limited amount of free iron available for this reaction to take place in vivo, multiple potential host and microbial catalytic iron complexes exist, as discussed previously. In vitro studies have demonstrated that increased bacterial iron concentrations enhance 'OH-mediated killing of S. aureus by H<sub>2</sub>O<sub>2</sub>, human monocytes, and neutrophil-derived cytoplasts (142, 256). However, the role of OH-mediated killing of S. aureus by intact human neutrophils remains unresolved (70, 257). In addition, killing of Leishmania donovani chagasi promastigotes by H<sub>2</sub>O<sub>2</sub> appears to involve iron-dependent OH formation (329), but these studies have not yet been extended to phagocyte systems. The role of iron in microbe-phagocyte interactions is clearly complicated, since Byrd and Horwitz have shown that conditions that modulate phagocyte iron concentration appear to affect intracellular microbicidal activity against Legionella pneumophila and M. tuberculosis in opposite ways (46, 47).

Recently, NO has been increasingly recognized as another phagocyte-derived oxidant involved in microbicidal activity. Its synthesis requires a NO synthase, of which there exist constitutive and inducible isoforms (see above) (216). The inducible enzyme has been definitively demonstrated in murine phagocytic cells and can be induced by a number of cytokines and lipopolysaccharides (216, 295). Despite efforts by many investigators, however, the ability to detect NO production by human mononuclear phagocytes has been modest at best under conditions where NO production by murine phagocytes is readily apparent (49, 117, 159, 160, 223). Recent data have demonstrated that human mononuclear phagocytes can produce constitutive NO synthase (255, 312). The inducible NO synthase mRNA and protein are generated in response to lipopolysaccharide and/or gamma interferon stimulation, demonstrating that human phagocytes appear to possess the necessary "machinery" to synthesize NO. More direct evidence for NO production by human macrophages has been demonstrated by the recent findings of Nicholson et al. (227). An average of 65% of alveolar macrophages in bronchoalveolar lavage specimens from 11 patients with untreated, culturepositive pulmonary tuberculosis contained NO synthase mRNA and functional NO synthase expression. Of note, only 10% of alveolar macrophages from normal subjects demon-

strated similar findings. However, despite these reports, the quantity of NO generated under a number of conditions was very small (312). In addition, studies where NO production is equated with nitrite production may falsely overestimate the true quantity of NO synthesis, as shown by Klebanoff and Nathan, who demonstrated that human neutrophils can synthesize nitrites via the catalase-catalyzed conversion of azide to nitrite in the presence of phagocyte-generated H<sub>2</sub>O<sub>2</sub> in vitro (168).

The primary microbicidal effect of phagocyte-derived NO appears to involve intracellular pathogens. A clear role in pyogenic bacterial infections has not been demonstrated. By treating murine-activated macrophages in vitro with  $N^G$ -monomethyl-L-arginine, a competitive inhibitor of nitrate and nitrite synthesis from L-arginine, a number of investigators have implicated NO as having microbiostatic and/or microbicidal activity against pathogens such as *Cryptococcus neoformans* (117), *Toxoplasma gondii* (1), *Mycobacterium bovis* (100), *Leishmania major* (180, 182), *Schistosoma mansoni* (160), and others (223, 261). Further studies involving an in vivo model of murine leishmaniasis have demonstrated that NO plays an important role in containing the extent of infection and decreasing the overall organism load (92, 181). A similar result was observed by Boockvar et al. in an in vivo model of murine listeriosis (28).

Although the importance of NO production to murine macrophage function is now well established, these data are not directly applicable to human phagocytes, because the existence of a role for NO in human phagocyte microbicidal activity is less clear. Recent data by Vouldoukis et al. suggest that the killing of *L. major* by human macrophages is mediated by NO, whose production is induced after cell activation via ligation of the low-affinity receptor for immunoglobulin E (FcERII/CD23 surface antigen) (307). This receptor is upregulated in cutaneous leishmaniasis. Additional in vitro data imply that tumor necrosis factor alpha (TNF- $\alpha$ ) and granulocyte-macrophage colony-stimulating factor stimulate human macrophages to restrict the growth of virulent Mycobacterium avium by a mechanism involving NO (80). The previously discussed findings by Nicholson et al. (227) also suggest that NO may be an important component of host defense against pulmonary tuberculosis. Likewise, Bukrinsky et al. reported that lipopolysaccharide or TNF-α-activated HIV-infected monocytes exhibit enhanced NO production (44). In support of these findings, the authors detected RNA encoding the inducible NO synthase in postmortem brain tissue from an AIDS patient with advanced HIV encephalitis. Nitric oxide may also contribute to the killing of staphylococci by neutrophil cytoplasts (anucleate, granulepoor, motile cells) which rapidly took up and killed the bacteria by a mechanism inhibited by N<sup>G</sup>-monomethyl-L-arginine (194). In contrast to these various studies, a direct comparative study between murine and human macrophages revealed that activated murine but not human macrophages demonstrated enhanced NO production and antimicrobial activity against Toxoplasma gondii, Chlamydia psittaci, and Leishmania donovani (223). Although not directly compared with murine macrophages, NO production contributes minimally to human macrophage-mediated killing of Cryptococcus neoformans and Schistosoma mansoni (49, 158), organisms which are killed by a NO-mediated mechanism in murine macrophages (117, 160). This suggests that NO makes a minimal contribution to the overall microbicidal activity against these pathogens in the human host.

Thus, there are increasing data supporting the concept that human phagocytes can produce NO, albeit in small quantities relative to their murine counterparts. However, these data are somewhat difficult to interpret, as the frequency with which negative findings are reported by laboratories is often quite low. This capability to synthesize NO appears to be mediated via the classic NOS pathway. The microbicidal activity of human, like murine, phagocyte-derived NO, if and when it is generated, could contribute to host protection against intracellular organisms. However, the contribution of NO relative to the other phagocyte antimicrobial mechanisms known to be effective against these and other pathogens has yet to be established.

Other oxidant sources and their contribution to microbicidal activity. Although phagocytes are the primary source of microorganism exposure to oxidants in mammalian hosts, other mechanisms of oxidant production exist and probably contribute to microbial oxidant stress. As discussed in a previous section, microorganisms such as Nocardia and Lactobacillus spp. produce NO and H2O2, respectively, which may in turn have toxic effects on other microorganisms in close proximity (64, 89, 167, 332). In addition, endothelial cells produce NO, O<sub>2</sub>, and H<sub>2</sub>O<sub>2</sub> in response to a number of stimuli, including inflammation (216). Feng et al. have recently suggested that endothelial cell-derived NO could protect these cells from infection with Rickettsia conorii (96). In an experimental model, pulmonary (tracheal and alveolar) epithelial cells also demonstrate luminal H<sub>2</sub>O<sub>2</sub> production, which is enhanced after stimulation by phorbol myristate acetate and platelet-activating factor (165). Epithelial cells from cystic fibrosis patients have been shown to consume two- to threefold more oxygen than do normal cells, providing indirect evidence of a highly oxidative environment in a population known to have a chronically high organism load (296). These endothelial cell- or epithelial cell-derived compounds may exert microbial oxidant stress either alone or via their reaction by-products such as ONOO and/or OH in the extracellular space. Additionally, these oxidants may interact with phagocyte-derived oxidants, cytokines, and other compounds to potentiate the microbial insult.

Several antimicrobial agents used in the treatment of clinical infections, in addition to blocking key enzymes and other metabolic functions of microorganisms, produce reactive oxygen intermediates that are capable of damaging other biomolecules. For example, β-lactam antibiotics (penicillins and cephalosporins) have been shown to oxidatively damage DNA and deoxyribose in the presence of iron and copper salts, consistent with an OH-mediated mechanism (246). In addition, the polyunsaturated structure of the polyene antifungal antibiotics (amphotericin, natamycin, and nystatin) gives them the propensity to oxidize to form peroxy radicals and thiobarbituric acid-reactive aldehyde fragments (281). These interactions can then lead to the generation of other oxygen-centered radicals capable of inciting further microbial injury. These newly recognized antibiotic effects may prove to be an important component of their biologic activities.

Likewise, a number of compounds undergo rapid redox cycling under aerobic conditions, potentially resulting in an additional source of extracellular oxidants for microbial encounter (135, 191). These compounds are univalently reduced to free radicals by cellular systems. In the presence of  $O_2$ , these reduced molecules are then reoxidized, with the resulting transfer of that electron to  $O_2$ , hence forming  $O_2^-$  and  $O_2^-$  the latter via  $O_2^-$  dismutation. Examples of such compounds include pharmacologic agents such as adriamycin, bleomycin, and nitrofurantoin (191).

The *P. aeruginosa* secretory product pyocyanin works by a similar mechanism. This compound is a phenazine-derived pigment that can undergo redox cycling to induce both intra- and

extracellular  $O_2^-$  and  $H_2O_2$  production from  $O_2$  in both eukaryotic and prokaryotic cells (134, 135, 218). This process contributes to cell death through the diversion of electron flow from normal biologic pathways into those leading to toxic oxidant generation. This pyocyanin-induced production of  $O_2^-$  and  $H_2O_2$  also can lead to the formation of  $O_2^-$  in the presence of a catalytic iron source (38, 208). Pyocyanin production increases under conditions of nutritional deprivation and oxidative stress (136). Interestingly, however, *P. aeruginosa* itself is relatively insensitive to pyocyanin and seemingly escapes oxidant-mediated injury during production of or exposure to this compound (136). This may be explained in part by its low endogenous levels of NADH/NADPH, its lack of NADPH:pyocyanin oxidoreductase, and/or its high levels of SOD and catalase.

In addition to its redox capabilities, pyocyanin has numerous in vivo and in vitro effects which could play a role in the pathogenesis of clinical infections. For example, the addition of pyocyanin, at concentrations detectable in pulmonary secretions of individuals with P. aeruginosa infection, to human ciliated nasal epithelial cells (4, 317, 318) and sheep tracheal epithelial cells (157) results in a loss of ciliary function as well as a decrease in in vivo tracheal mucus velocity in the sheep model (222). The effect on sheep cilia could be negated by the simultaneous presence of catalase, suggesting an oxidant-mediated mechanism (222). This process may contribute to the difficulty that cystic fibrosis patients experience in mobilizing their secretions (222). Other effects of pyocyanin potentially relevant to microbial killing and inflammatory tissue injury include those on stimulated neutrophils to alter O<sub>2</sub> production and degranulation, host cell NO production, and lymphocyte proliferation and differentiation (207, 219, 220, 230, 282, 305, 310).

# MECHANISMS OF MICROBIAL DEFENSE AGAINST OXIDANTS

#### Avoidance of Encounters with Phagocyte-Derived Oxidants

As microbial killing by phagocytes is a multistep process, microorganisms have likewise developed a sequential series of defense strategies to counteract this process. Some microorganisms secrete toxins to kill the phagocyte before they can be killed by it. Examples include the production of streptolysin by Streptococcus spp. (22), leukocidin by Staphylococcus aureus (263), and the *Clostridium septicum* toxin (193). Other organisms resist phagocytic uptake by covering their surfaces with hydrophobic capsules (Neisseria meningitidis, pneumococci) (103, 141, 147, 158, 258–260, 324). Pathogenic mucoid strains of P. aeruginosa synthesize alginate, an exopolysaccharide. In addition to aiding in avoiding phagocytic uptake, alginate has the ability to scavenge reactive oxygen intermediates, suppress leukocyte function, and promote bacterial adhesion (11, 82, 87, 177, 239, 280). This may be of particular clinical relevance, as airway isolates from individuals with cystic fibrosis commonly demonstrate alginate production (82, 239). P. aeruginosa also requires a unique glucose-dependent pathway for phagocytosis by macrophages (13). This may enhance its pathogenicity in the bronchoalveolar space, where concentrations of glucose and other carbohydrate are low. Acidification within the phagocytic vacuole is an important process to maximize the spontaneous dismutation of  $O_2^-$ , hydrolase activity, and phagosome-lysosome fusions. Inhibition of this acidification process has been described for Legionella pneumophila (146) and Toxoplasma gondii (279). Although phagocyte-derived oxidants are important mediators in microbial killing, some organisms

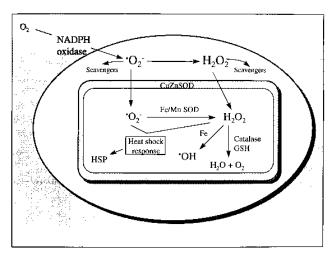


FIG. 2. Overview of bacterial defense mechanisms against oxidative killing inside the phagosome.

can survive the encounter to then inhibit phagosome-lysosome fusion and avoid enzymatic attack by hydrolytic enzymes. This process is poorly understood but has been demonstrated among some mycobacteria (143), *T. gondii* (143), *Chlamydia* spp. (143), and others (41, 146). Other organisms, such as *Listeria monocytogenes* (60), *Shigella flexneri* (274), and *Trypanosoma cruzi* (5), are able to escape from the phagosome by the secretion of membrane-damaging cytolysins.

Another key virulence factor allowing for the avoidance of host defense mechanisms has been identified in a number of Yersinia spp. In Yersinia enterocolitica, a 51-kDa periplasmic protein encoded by the yop-51 gene shares amino acid sequence identity with the catalytic domain of several protein tyrosine phosphatases (PTPases) (120). Activation of protein tyrosine kinases is an important signaling mechanism in many cells, including macrophages. By interfering with host signaling pathways, Yersinia spp. have the potential to modify the host immune response, which probably explains the importance of this process as a virulence factor. The yop-51 gene resides on a naturally occurring 70-kb plasmid, and its mutation alters the virulence of the organism (26). An analogous gene, *yopH*, encodes a similar PTPase in *Y. pseudotuberculosis*. Further work has characterized the crystalline structure and active site of these proteins (294, 331). The specific gene and corresponding PTPase have not been determined for Y. pestis, but preliminary studies reveal that the yop-51/yopH gene is highly conserved in this organism (242).

# **Defense Strategies Specific for Oxidants**

Nonenzymatic. Exposure to intraphagosomal oxidants is a fatal event for many microorganisms. However, some organisms have evolved an ability to inhibit the NADPH- oxidase-dependent oxidative burst and thus to inhibit reactive oxidant production within the phagosome (Fig. 2). This appears to be particularly important for intracellular pathogens as it aids in survival within the phagosome. For example, the lipophosphoglycan present on the membrane of *Leishmania major* and *L. donovani* (analogous to lipopolysaccharide in bacteria) inhibits protein kinase C activity in macrophages (30, 104), resulting in suppression of the respiratory burst and ultimately of 'O<sub>2</sub><sup>-</sup> production. This inhibition of macrophage protein kinase C activity also impedes macrophage chemotactic locomotion and interleukin-1 (IL-1) production. *Legionella pneumophila* se-

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cretes a compound shown to inhibit the neutrophil oxidative burst (272). *Leishmania donovani* (114) and *Legionella micdadei* (271) produce extracellular acidic phosphatases that block  $O_2^-$  formation in vitro. The mechanisms of these effects have not been further elucidated, however.

Antioxidant scavengers unique to specific pathogens have also evolved to protect microorganisms from phagocyte-derived oxidants. As noted above, P. aeruginosa produces alginate, an exopolysaccharide capable of scavenging oxidants (177, 280). In addition, the phenolic glycolipid of mycobacteria and the lipophosphoglycan of L. donovani are effective in scavenging 'OH and 'O<sub>2</sub><sup>-</sup>; these characteristics may enhance the intracellular survival of the organisms (61, 226). Cryptococcus neoformans is known to produce large amounts of mannitol both in vivo and in vitro (323). Mannitol in high concentrations has the ability to scavenge reactive oxygen species. Thus, its production by C. neoformans may be a protective mechanism by which the organism protects itself from oxidative killing by host phagocytes. Chaturvedi et al. have recently demonstrated, by using a low-mannitol-producing mutant of C. neoformans, that the ability of C. neoformans to produce and accumulate mannitol may influence its tolerance to heat and osmotic stresses and its pathogenicity in mice (62) through the scavenging of reactive oxygen intermediates (63).

The formation of heat shock proteins (HSP) by ingested bacteria may represent another adaptive mechanism. HSP production can be induced by increased temperature and/or oxidant exposure as a means of protection against both heat and oxidant damage. In *Mycobacterium tuberculosis* and *Mycobacterium leprae*, a strongly immunogenic antigen can be recognized by use of monoclonal antibodies (327). Production of this protein can be induced by stress, which may include phagocytosis. Severe stresses also increase the production of antioxidant enzymes such as SOD. There exists some evidence that HSP may also play a role in the regulation of antioxidant enzyme production in *E. coli* (293); this is discussed in more detail below.

Little is known about microbial defense against NO. During the process of denitrification, microorganisms appear to limit toxicity by keeping endogenous NO levels very low (334). In an in vitro model, extracellularly generated NO was inactivated by the *P. aeruginosa*-derived phenazine pigment pyocyanin (310). Once phagocytized, microorganisms may have evolved a strategy to inhibit host nitric oxide synthase analogous to what has evolved for the NADPH-oxidase complex. However, at present, this has not been reported.

**Enzymatic.** Microorganisms have developed highly specific and effective enzymatic pathways of oxidant inactivation, including those catalyzed by SOD, catalase/peroxidase, and glutathione (GSH) in combination with glutathione peroxidase and glutathione reductase (122, 137). (See Fig. 2 and Table 1 for chemical reactions.)

Glutathione serves as a substrate for the  $\rm H_2O_2$ -removing enzyme glutathione peroxidase. It can then be redox cycled via glutathione reductase for further  $\rm H_2O_2$  removal. GSH is also an OH scavenger. Eukaryotic cells depleted of GSH exhibit increased susceptibility to oxidant-mediated killing (205). There are also data suggesting that GSH depletion is involved in HIV replication (162, 286). The importance of this antioxidant system in prokaryotes, however, has not been clearly established. GSH reductase-negative *E. coli* mutants do not demonstrate an increased susceptibility to  $\rm H_2O_2$ -mediated stress compared with the isogenic parental strain (118). However, there are data suggesting that GSH may facilitate the deactivation of *E. coli* aconitase and other [4Fe-4S]-containing dehydratases that have been oxidatively inactivated by  $\rm O_2^-$ 

(111). Proteins immunologically related to GSH have been demonstrated in other bacterial species and in other strains of E. coli (240). Recently, Moore and Sparling have identified a GSH peroxidase homolog gene, gpxA, in Neisseria meningitidis. The amino acid sequence of this gene is highly homologous to GSH peroxidases found in other bacterial species (217). Thus, there may be several types of GSH-metabolizing proteins in bacteria which serve a similar purpose, and their distribution may even vary within a single species. Protozoa such as trypanosomes and leishmaniae produce trypanothione (93). It may have an analogous function to GSH in that it functions to maintain thiol redox within the organism and as a defense mechanism against oxidants, xenobiotics, and heavy metals. The importance of trypanothione to parasite survival can be exemplified by organism exposure to D,L-α-difluoromethylornithine, an antiparasitic agent used in the treatment of human African trypanosomiasis. D,L-α-Difluoromethylornithine inhibits parasite ornithine decarboxylase, which results in decreased cellular trypanothione levels among other effects (93). Why these organisms have evolved to produce trypanothione in addition to GSH is unclear.

Considerably more data are available on the distribution, structure, and regulation of microbial catalases and peroxidases (186). The antioxidant action of these enzymes is to catalytically convert  $H_2O_2$  to  $H_2O$  and  $O_2$ . Nearly all aerobic and facultatively anaerobic microorganisms, with the exception of the *Streptococcus* spp., synthesize at least one form of catalase and/or peroxidase (201). The majority of obligate anaerobes lack this capability (201). These proteins are characterized by structural diversity between different organisms and even within the same organism. The most common form consists of a homotetramer with one protoheme IX per subunit. Most bacteria produce two catalases, whereas others such as *Klebsiella pneumoniae* and *P. aeruginosa* have the ability to produce multiple catalases under specific growth conditions (115, 136).

The two structurally distinct catalases of E. coli, termed hydroperoxidase I (HPI) and hydroperoxidase II (HPII), have been the most extensively studied (65, 66). HPI, a bifunctional catalase-peroxidase encoded by *katG*, contains two protoheme IX groups associated with a tetramer of identical 80-kDa subunits and is localized in the periplasmic space. HPII, a monofunctional catalase encoded by katE, consists of six heme d isomers associated with a hexameric structure of 84.2-kDa subunits and is found solely in the cytoplasm. The relative levels of HPI and HPII are controlled by two different regulons that respond to different environmental stimuli (186). HPI is synthesized preferentially in response to oxidative stress (H<sub>2</sub>O<sub>2</sub>), whereas HPII is produced in response to nutrient depletion as occurs in the stationary growth phase (188). Thus, not only are HPI and HPII different structurally and genetically, but also the processes controlling their synthesis respond to different stimuli and involve different mechanisms. The two catalases of Bacillus subtilis have been studied in comparison and appear to show some resemblance to E. coli HPI and HPII with regard to their structure and mechanism of control (185, 187). Among other bacterial species, the catalases of several other members of the Enterobacteriaceae family exhibit homology to E. coli HPI and HPII (298).

The complexity of bacterial catalase expression and regulation can be demonstrated by the reported correlation between the loss of catalase production and isoniazid resistance among *Mycobacterium tuberculosis* isolates (98, 330). Different amounts of catalase production have been found in a number of organisms in response to nutrient depletion and in association with their susceptibility to phagocyte killing (136, 188,

197). For example, the growth of *P. aeruginosa* in limited-succinate media resulted in increased catalase activity and the appearance of additional catalase isoforms compared to the catalase activity in the same organisms grown under nutritionally replete conditions (136, 210). Mandell demonstrated that neutrophils easily killed low- but not high-catalase-producing *Staphylococcus aureus* strains (197). This difference correlated with *in vivo* lethality in a mouse model. Likewise, catalase-deficient *E. coli* mutants exhibit an increased susceptibility to phagocyte-mediated killing (121).

Another mechanism of oxidant inactivation used by microorganisms involves SOD. The production of this group of enzymes is a key defense strategy aimed at the elimination of O<sub>2</sub>-. Not only does this decrease the possibility of direct O2-mediated toxicity, but also it prevents O2-mediated reduction of iron and subsequent 'OH generation via the Haber-Weiss reaction. There are three common forms of SOD found in nature (12). Eukaryotes and some higher fungi predominantly produce CuZnSODs, homodimers (molecular weight 32,000) with two noncovalently linked identical subunits containing one atom each of copper and zinc. A few species of bacteria have also been found to contain CuZnSOD; some of these include Stenotrophomonas (Pseudomonas) maltophila (289), Brucella abortus (18), several Haemophilus spp. (171, 172, 176), E. coli (21), N. meningitidis (173), L. pneumophila (291), and Salmonella spp. (52). All bacteria, including obligate anaerobes, produce either FeSOD, MnSOD, or both. Like CuZnSOD, these enzymes exist as subunits (molecular weight 23,000) linked as dimers in FeSOD and dimers and tetramers in MnSOD. The metal content of both isozymes varies between 1 and 2 atoms per dimer. Most SODs are cytoplasmically located, although a few are located on or secreted through the cytoplasmic membrane (16, 21, 270). In general, FeSOD predominates in anaerobic organisms whereas MnSOD is more commonly found among aerobic organisms. Although variations in the FeSOD content have been observed in bacteria producing both isoenzymes, it is the control of MnSOD that is usually responsible for modulating the total level of SOD in

Like catalase, microbial SOD regulation and genetics have been most extensively studied in E. coli (302), where SOD expression is dependent on a number of environmental stimuli. FeSOD, encoded by the *sodB* gene, is produced constitutively in E. coli grown aerobically or anaerobically but is upregulated when grown anaerobically in the presence of nitrate. MnSOD becomes the predominant form when the cell is exposed to oxidative stress. In addition, a hybrid form containing Fe and Mn has also been isolated in vitro (85). It appears that, functionally, FeSOD provides E. coli with the first line of defense against 'O<sub>2</sub> and MnSOD is subsequently recruited in circumstances of increased oxidative stress. MnSOD gene (sodA) expression is governed by a number of regulon proteins, such as the Fur proteins and those under the control of the sox gene locus, including the Arc protein (119, 228, 304). These proteins are made in response to such stimuli as iron availability and oxygen/oxidant exposure, respectively. Interestingly, the Arc regulatory protein is also involved in the control of aconitase synthesis, suggesting that increased MnSOD is necessary to protect increased cellular concentrations of aconitase (109). Evidence suggests that MnSOD regulation also occurs at the posttranscriptional and posttranslational levels (244).

In recent work, Hassett et al. have begun to characterize a similar system in *P. aeruginosa*, which, like *E. coli*, possesses both an iron- and manganese-cofactored SOD (136, 138). Notably, *P. aeruginosa* has approximately four to five times the SOD activity reported for *E. coli* (136). When cloned, the

genes encoding the MnSOD (sodA) and FeSOD (sodB) revealed a 50 and 67% sequence homology with the respective E. coli SODs. The relative quantities of the FeSOD or MnSOD isoenzyme produced appear, as in E. coli, to be dependent on nutritional availability and the degree of oxidative stress (136).

Some pathogens have evolved the ability to localize SOD activity to their extracellular environment as a means of resisting oxidant attack. For example, Mycobacterium tuberculosis secretes an FeSOD, whereas Legionella pneumophila and E. coli possess a periplasmic CuZnSOD (21, 270). Nocardia asteroides has a unique SOD associated with the outer cell wall, which can be selectively secreted extracellularly. This SOD differs significantly from those isolated from other bacteria in that it contains equimolar amounts of Fe, Mn, and Zn (16). Although N. asteroides induces an oxidative burst in human phagocytes, it is not readily killed by this mechanism. Subsequent in vitro and in vivo studies have demonstrated that this resistance to phagocyte-mediated killing is dependent on the production and secretion of SOD by the organism (15, 17). Initial killing and/or enhanced clearance of N. asteroides was observed in organs obtained from infected mice given a monoclonal anti-SOD antibody-treated N. asteroides. This effect was not observed in mice given a nonspecific nocardial antibody. Thus, the extracellular localization of bacterial SOD may be an important determinant in the pathogenesis of infection for N. asteroides and other pathogens.

The importance of microbial SOD production can be appreciated when studying SOD-deficient organisms. MnSOD- and FeSOD-negative mutants have been obtained from  $E.\ coli$  (303). However, with the recent discovery of the periplasmic  $E.\ coli$  CuZnSOD (21), studies with these mutants warrant qualification. Nonetheless, these organisms exhibit extreme sensitivity to oxidizing agents such as paraquat and methylene blue and are more susceptible to phagocyte-mediated killing (53, 121). Fe- and Mn-SOD-deficient double mutants demonstrate a marked increase in oxygen-dependent mutagenesis (94). Amino acid biosynthesis and membrane integrity also appear to be affected (151, 152). However, in vitro data obtained with  $E.\ coli$  suggest that overexpression of SOD may also be deleterious by accelerating  $H_2O_2$  production in the organism upon its exposure to oxidative stress (276).

# ROLE OF OXIDANTS IN VIRAL INFECTIONS

There is recent evidence that oxidants, whether derived from phagocytes or other sources, play a role in the pathogenesis of viral infections. The majority of work has centered around HIV. HIV infection is associated with a proinflammatory state in the host, resulting in high levels of circulating cytokines, including TNF-α, IL-1α, IL-1β, IL-2, IL-6, alpha interferon (IFN- $\alpha$ ), and IFN- $\gamma$  (262). Although it has been shown that some of these cytokines can activate HIV replication in the infected host cell directly (241), cytokine activation of phagocytes and other cells can also stimulate oxidant production. Oxidants also have direct effects on HIV replication. Legrand-Poels et al. demonstrated that the addition of exogenous H<sub>2</sub>O<sub>2</sub> to a latently HIV-infected T-cell line (U1) resulted in increased replication of the HIV-1 provirus (178). Schreck et al. confirmed these findings in Jurkat T cells and provided further insight into the mechanism of activation (275). This process, like direct cytokine activation of HIV, is mediated by the induction of NF-kB, a ubiquitous transcription factor that is recognized by the HIV promoter (275). Likewise, Sandstrom et al. showed that HIV gene expression enhanced T-cell susceptibility to H<sub>2</sub>O<sub>2</sub>-induced apoptosis (273).

HIV-infected cells may be uniquely sensitive to oxidant

stress, as a number of studies have shown them to exhibit low levels of GSH, the main intracellular defense against oxidants. HIV-infected patients demonstrate decreased GSH levels in blood and peripheral blood mononuclear cells relative to those in normal patients, and this decrease becomes more pronounced with advanced disease (81). More specifically, Staal et al. found that in patients with symptomatic AIDS, GSH concentrations in CD8 and CD4 T cells are 62 and 63%, respectively, of those found in seronegative controls (286). The greatest decreases in GSH levels were seen in those patients with advanced infection. Not only does this decrease in intracellular GSH levels leave the infected cell susceptible to the direct effects of oxidants, but also it leads to increased NF-kB expression, resulting in further activation of HIV replication (285). Using a HeLa cell line transfected with the tat gene from HIV-1, Flores et al. found that the expression of the regulatory Tat protein, essential for virus replication, suppresses the expression of cellular MnSOD (102). These cells also exhibited other evidence of increased oxidative stress manifested by elevated levels of carbonyl proteins and decreased cellular sulfhydryl content (102). Thus, HIV-mediated modification of host antioxidant enzymes may be an important component in mediating ongoing HIV infection and the ultimate progression to severe immunodeficiency. This process may be further altered in the presence of opportunistic pathogens.

These advances in the understanding of the pathogenesis of HIV infection have prompted investigations into the use of antioxidants as therapy for HIV-infected individuals. In vitro studies with an HIV-infected human promonocytic cell line have demonstrated that HIV expression can be decreased by treatment of the cells with GSH, glutathione ester, or N-acetylcysteine (162, 195). Each of these compounds increases intracellular thiol concentrations and, as a result, inhibits NF-kB and ultimately HIV expression. These observations have led to studies of HIV-infected patients to determine whether the administration of N-acetylcysteine or L-2-oxothiazolidine-4carboxylic acid (Procysteine) may alter disease progression (81, 161). Although both of these compounds were found to increase intracellular GSH levels in treated patients, there were no significant differences in CD4 cell counts, viral load, or proviral DNA frequency. Additional in vitro data suggest that the oxidant scavenger ascorbate also suppresses HIV replication in chronically and acutely infected T cells (129). This interaction appears to be synergistic when cells are exposed to ascorbate and N-acetylcysteine concurrently (128). In vivo studies with this combination have not been reported to date. Although oxidants may play a role in the pathogenesis of HIV infection, applying these findings for the development of potential therapeutic strategies in HIV-infected patients has been of limited benefit thus far.

Oxidants may also be involved in the pathogenesis of other viral infections. As in HIV-infected cells, H<sub>2</sub>O<sub>2</sub> effectively induces synthesis of viral antigens in several lymphoid cell lines that harbor the Epstein-Barr virus genome (234). In contrast, H<sub>2</sub>O<sub>2</sub> markedly decreases the release of progeny hepatitis B virus (HBV) particles in cultured hepatoma cells without causing any significant difference in the overall pattern of host protein synthesis (333). These findings may be important in the pathophysiology of chronic HBV infection. In one circumstance, turning off viral gene expression may be a way for the host to eradicate HBV infection. However, this mechanism may allow the virus to evade complete destruction by shutting off viral expression in infected hepatocytes adjacent to an area of active inflammation. This would allow a few cells to escape antigen-specific killing and resume viral replication once the inflammation subsides. Levels of vitamin E in plasma are notably low in patients with chronic liver disease (306). It is widely established that vitamin E is an important cell membrane antioxidant which acts as a free radical scavenger. One might speculate that its deficiency in this setting may further perpetuate tissue damage caused by oxidant release from injured hepatocytes in patients with chronic viral hepatitis and ultimately with cirrhosis. However, there have been no controlled clinical trials assessing the therapeutic role of vitamin E in these patients.

Human papillomavirus infection has been linked to an increased risk of acquiring human cervical carcinoma, and a recent study by Fernandez et al. suggests a potential oxidantdependent mechanism which could be involved (97). They demonstrate that approximately 50% of healthy women possess polyamine oxidase and/or diamine oxidase in their cervical mucus. These enzymes were shown to act on spermine and spermidine (polyamines present in seminal fluid) to generate H<sub>2</sub>O<sub>2</sub> and reactive aldehydes, which are likely to exert local mutagenic effects in vivo. These transformed cervical cells may exhibit prolonged survival in the presence of HPV infection through HPV suppression of apoptosis in the keratinocytes. Thus, the authors suggest that the effects of HPV infection of cervical cells may be synergistic with the effects of polyamine oxidation occurring in the cervical environment of sexually active women. The regulation of HPV replication may also be modified by oxidants, as the intracellular redox environment has been shown to affect the posttranslational DNA-binding activity of three E2 proteins (199).

Virus-host cell interactions in relation to oxidant production also appear to be important in the pathogenesis of influenza A virus infection. Although neutrophils predominate in the early inflammatory response to influenza A virus (106), the ability of this virus to adversely affect neutrophil and monocyte function in infected patients is well established (133) and may contribute to secondary bacterial infections. The influenza A virus hemagglutinin molecule appears to be an important mediator in this process of abnormal leukocyte function (56, 131). Although exposure to the virus leads to neutrophil activation and generation of a respiratory burst, the neutrophil response is atypical with regard to calcium fluxes, phospholipase C activation, and release of H<sub>2</sub>O<sub>2</sub> but not 'O<sub>2</sub><sup>-</sup> (130, 132). Daigneault et al. have further characterized this unique virus-phagocyte interplay, specifically through studies of the hemagglutininneutrophil receptor interaction (78). Clearly, further understanding of the role of oxidants in viral replication and virushost cell interactions for these and other viruses could potentially lead to new therapeutic interventions.

# UNTOWARD CONSEQUENCES OF OXIDANT PRODUCTION FOR THE HOST

At sites of infection, host-derived oxidants not only place the offending organisms under oxidative stress but also cause stress to neighboring host tissues. As discussed above, these oxidants are derived primarily from phagocytes; however, they can be produced by other cell types inherently or via induction by redox-active agents. Tissue injury at sites of infection may be the result of the host inflammatory response to the pathogen rather than cytotoxic components of the microorganism. The role of oxidants in such processes will briefly be reviewed, given their intimate relationship with the pathophysiology of many infectious diseases. Readers are referred to the myriad of excellent recent reviews on oxidant-mediated tissue injury (76, 124–126, 209, 216, 309).

Many aspects of acute and chronic inflammatory tissue injury appear to be mediated by oxidants released by neutrophils

and other phagocytes (216, 313). This process is enhanced by adherence of the phagocyte to the target cell surface (309). This adherence and subsequent movement of phagocytes from the blood to sites of inflammation require a complex signaling system involving a family of glycoproteins termed selectins. Selectins are synthesized by endothelial cells and stored in their secretory granules. When endothelial cells are activated by compounds such as thrombin or histamine (released in response to inflammation), the granules fuse with the outer membrane to expose the selectins on the cell surface. Phagocytes recognize these proteins, and this promotes their adherence to the endothelium and primes them for degranulation (190, 203). Following leukocyte activation, phagocyte-derived proteins termed integrins bind to their respective receptors on the endothelial cell. This interaction further strengthens adhesion and directs the migration of the phagocyte beneath the endothelium. Thus, this process targets phagocytes to areas of inflammation, where, through the further recruitment of phagocytes, oxidant-mediated tissue injury may result. Phagocyte-derived H<sub>2</sub>O<sub>2</sub> can also indirectly lead to inflammatory tissue injury by upregulating selectin expression on endothelial cells and promoting further neutrophil localization (238).

Inflammatory tissue injury may also result via oxidant-induced cellular production of proinflammatory cytokines (72, 164, 206). Likewise, the production of these cytokines may potentiate further cellular oxidant release. Many of these interactions are mediated through the transcription-regulating factor, NF-κB (266). For example, in a rat model of neutrophilic alveolitis, endotoxin-induced NF-kB activation is thought to mediate the production of cytokine-induced neutrophil chemoattractant (analogous to human IL-8) by alveolar macrophages (25). This process is believed to be important for the recruitment of neutrophils and ultimately for the inflammatory tissue injury seen in this model. Oxidants can also activate NF-κB, promoting the production and release of cytokines such as IL-1 and TNF- $\alpha$  (8). Joint inflammation can also be induced by bacterial products, immune complexes, and crystals which recruit and activate phagocytes to form reactive oxygen species (125, 126) primarily. This process can result in tissue destruction via oxidant interactions with host proteoglycans, collagen, and elastin (124).

Injury to pulmonary epithelial and pulmonary vascular endothelial cells can also occur as a consequence of microbial infection in the case of acute necrotizing pneumonia and chronic lung infection seen in cystic fibrosis patients (278, 309). A similar injury pattern can be observed with infection-related pulmonary complications such as acute respiratory distress syndrome and hyperoxic lung injury. The principal mechanism by which this lung injury occurs remains to be determined, but it appears to involve alterations in a number of parameters of epithelial and endothelial cell function inducible by phagocytederived  $O_2^-$  and/or  $H_2O_2$  (278). Iron-dependent formation of OH appears to be involved in the ability of phagocytes to damage endothelial cells in vitro, with the endothelial cells serving as the source of catalytic iron (107, 175). Peroxynitrite also has been shown to inhibit pulmonary epithelial cell ion channels, suggesting that this species could contribute to diffusion barrier disruption under conditions in which both 'O<sub>2</sub> and NO are present concurrently (14).

MPO-derived oxidants released in response to a microbial stimulus may also contribute to inflammatory tissue injury directly via their toxic effects (148) and indirectly by their ability to inactivate serine protease inhibitors such as  $\alpha_1$ -antitrypsin (54, 83). These antiproteases play a critical role in limiting the activity of proteases such as human neutrophil elastase released at local sites of inflammation (202). Thus, protease

inhibitor inactivation by MPO-derived oxidants may lead to emphysematous changes analogous to those seen in individuals congenitally deficient in  $\alpha_1$ -antitrypsin. Such processes have been hypothesized to contribute to lung injury associated with chronic bronchitis and other forms of chronic obstructive pulmonary disease (198, 233, 315). Others have suggested this process may also be involved in the lung disease observed in cystic fibrosis patients (214, 297).

Data supporting a role for NO and its derivatives in mediating inflammatory tissue injury in humans have been limited mainly to studies of autoimmune diseases (95, 216, 288). Evidence supporting NO production at sites of infection is lacking, however, as there are no definitive data demonstrating its formation by human phagocytes in vivo. In fact, recent literature suggests that NO may also have antioxidant properties (163, 288, 319). However, bacterially derived lipopolysaccharide induces NO production in endothelial cells. This process may contribute to the vasodilation and hypotension observed in septic shock (216).

Like microorganisms, host cells have evolved a complex system to defend themselves against oxidant injury. As discussed above, eukaryotic cells synthesize CuZnSOD as a means of  $O_2$  elimination. This enzyme is located in the cytosol and is usually constitutively expressed (12). Synthesis of a manganese-containing enzyme (MnSOD) can also be induced in the mitochondrial matrix under conditions of increased oxidative stress, specific cytokine stimulation, or heat shock (12). Since the  $H_2O_2$  formed by the dismutation of  $O_2^-$  is also cytotoxic, eukaryotic cells have developed various mechanisms for its removal analogous to those found in prokaryotic microorganisms. This is accomplished by regulation of intracellular levels of catalase, the two GSH-dependent enzymes, GSH, and/or NADPH. Intra- and extracellular oxidant scavengers, such as ascorbic acid, vitamin E,  $\beta$ -carotene, and  $\alpha$ -tocopherol, also probably play an important role in limiting cellular susceptibility to oxidant-mediated injury (313). Preventing the formation of  $O_2^-$  and  $H_2O_2$  is the primary mechanism by which cells can limit the formation of other potent oxidants such as OH and the MPO-derived oxidants. Hydroxyl radical generation via the Haber-Weiss reaction can also be controlled by limiting the availability of redox-active iron catalysts through the formation of less active iron complexes such as extracellular lactoferrin and transferrin (7, 9, 42, 45, 320) and intracellular ferritin (10, 59). Heme oxygenase mRNA expression in mammalian cells is also known to be increased following cell exposure to oxidant stress. Although disputed by some investigators (231), the proposed mechanisms of protection afforded by heme oxygenase induction are twofold (6, 292). Heme oxygenase decreases the availability of intracellular iron capable of participating in the Haber-Weiss reaction by catalyzing the conversion of free heme to bile pigments. These bile pigments in turn exert antioxidant effects. Little is known about how host cells protect themselves from injury by NO. It is likely that regulation of its production by the cell-specific NO synthase will prove important.

The extent of oxidant-mediated cytotoxicity observed at sites of inflammation is dependent on the balance between host-and microorganism-derived prooxidant and antioxidant forces. When this balance is swayed in favor of the prooxidants, not only microbial but also host cell cytotoxicity results, leading to clinical manifestations such as the sepsis syndrome, acute respiratory disease syndrome, lung destruction in diseases such as cystic fibrosis and  $\alpha_1$ -antitrypsin deficiency, and joint destruction in inflammatory arthritides. Further understanding of the mechanisms that regulate the prooxidant-antioxidant balance will probably have significant therapeutic implications in

the management of these and other diseases characterized by inflammatory tissue injury.

#### **CONCLUSIONS**

Defining the many roles of reactive oxygen species in hostmicrobial interactions has proved complex. Although these oxidants are consistent by-products of normal cellular metabolism, the concentration and potential biotoxicity can be markedly enhanced under conditions of exogenous oxidative stress, by exposure to pharmacologic agents, and, particularly, by phagocytes as a means of host defense against invading microorganisms. These oxidants can have beneficial and detrimental functions in both the host and the microorganism. Therefore, both have evolved complex adaptive mechanisms for protection against these compounds, including enzymatic and nonenzymatic oxidant-scavenging systems. These systems act as virulence factors for the microorganism which enable it to survive in a hostile environment. Despite the marked progress in this field recently, there are still many unanswered questions regarding the role of oxidants in microbial pathophysiology that will probably prove to be a promising research area in the future.

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